

CHEMISTRY

Paper 1 Multiple Choice

0620/13

May/June 2011 45 Minutes

Additional Materials:	Multiple Choice Answer Sheet
	Soft clean eraser
	Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers A, B, C and D.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

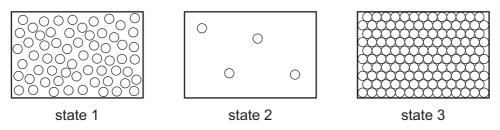
Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer. Any rough working should be done in this booklet. A copy of the Periodic Table is printed on page 16. You may use a calculator.

This document consists of **16** printed pages.



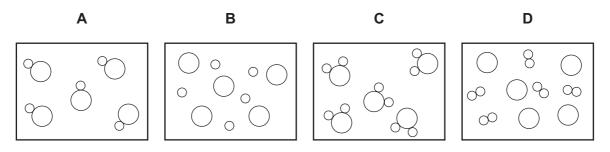
1 The diagrams show the arrangement of particles in three different physical states of substance X.



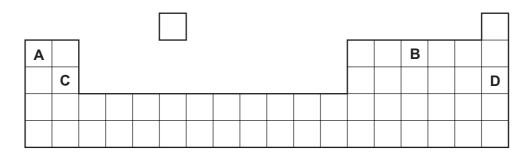
Which statement about the physical states of substance X is correct?

- A Particles in state 1 vibrate about fixed positions.
- **B** State 1 changes to state 2 by diffusion.
- **C** State 2 changes directly to state 3 by condensation.
- **D** The substance in stage 3 has a fixed volume.
- 2 In the diagrams, circles of different sizes represent atoms of different elements.

Which diagram represents hydrogen chloride gas?



3 The diagram shows part of the Periodic Table.



Which element is correctly matched with its electronic structure?

	electronic structure
Α	2,8,1
в	2,4
С	2,8,2
D	2,8

4 An aqueous solution is coloured.

Which method of separation would show that the solution contains ions of different colours?

- **A** chromatography
- **B** crystallisation
- **C** distillation
- **D** filtration
- 5 The table gives the solubility of four substances in ethanol and in water.

A mixture containing all four substances is added to ethanol, stirred and filtered.

The solid residue is added to water, stirred and filtered.

The filtrate is evaporated to dryness, leaving a white solid.

Which is the white solid?

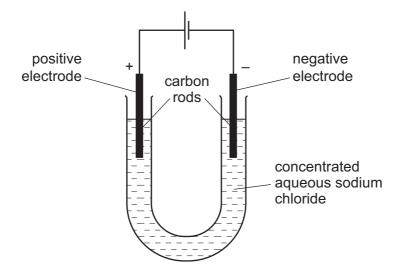
	solubility in					
	ethanol water					
Α	insoluble insoluble					
в	insoluble soluble					
С	soluble insoluble					
D	soluble	soluble				

6 Which two elements react together to form an ionic compound?

	element	electronic structure	
	W	2,4	
	x	2,8	
	Y	2,8,1	
	Z	2,8,7	
в	X and Y C	Y and Z D Z	Z and W

A W and X

7 Electricity is passed through concentrated aqueous sodium chloride, as shown.



What is the test for the gas formed at the positive electrode?

- A bleaches damp litmus paper
- **B** 'pops' with a lighted splint
- **C** relights a glowing splint
- **D** turns damp red litmus paper blue
- 8 Electricity from a power station passes through overhead cables to a substation and then to a school where it is used to electrolyse concentrated hydrochloric acid using inert electrodes.

Which substances are used for the overhead cables and for the electrodes?

	overhead cables	electrodes
Α	aluminium	copper
В	aluminium	platinum
С	copper	platinum
D	platinum	aluminium

9 The nucleon number and proton number of the lithium atom are shown by the symbol ${}_{3}^{7}$ Li.

What is the correct symbol for the lithium ion in lithium chloride?

 $\textbf{A} \quad {}^{6}_{2} Li^{-} \qquad \textbf{B} \quad {}^{6}_{3} Li^{+} \qquad \textbf{C} \quad {}^{7}_{3} Li^{+} \qquad \textbf{D} \quad {}^{7}_{3} Li^{-}$

burning methane in air

radioactive decay of ²³⁵U

reacting hydrogen with oxygen.

Which statements about these processes are correct?

- 1 Hydrogen and methane are being used as fuels.
- 2 All the processes involve oxidation.
- 3 All the processes are used to produce energy.

A 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3

- 11 Which statement about the electrolysis of molten lead(II) bromide is correct?
 - **A** A colourless gas is seen at the cathode.
 - **B** A grey metal is seen at the anode.
 - **C** A red/brown gas is seen at the anode.
 - **D** A red/brown metal is seen at the cathode.
- **12** What is the relative molecular mass (M_r) of HNO₃?
 - **A** 5 **B** 31 **C** 32 **D** 63
- **13** The equation for the effect of heat on hydrated sodium carbonate is as shown.

 $Na_2CO_3.10H_2O(s) \rightleftharpoons Na_2CO_3(s) + 10H_2O(g)$

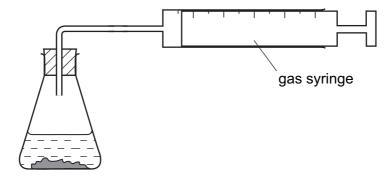
Statements made by four students about the reaction are given.

- P Anhydrous sodium carbonate is formed.
- **Q** Steam is formed.
- **R** There is a colour change from blue to white.
- **S** The reaction is reversible.

Which students' statements are correct?

- A P, Q and R only
- **B** P, Q and S only
- **C** Q, R and S only
- D P, Q, R and S

14 The apparatus shown can be used to measure the rate of some chemical reactions.

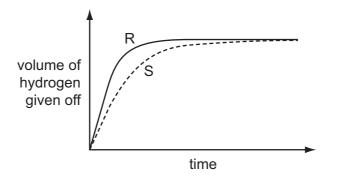


For which two reactions would the apparatus be suitable?

1 and 2 B	1 and 3 C	2 and 4	D	3 and 4
reaction 4	$ZnCO_3(s) + 2HCl(aq) \rightarrow ZnCl_2(aq) + CO_2(g) + H_2O$		$O_2(g) + H_2O(I)$	
reaction 3	$MgO(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2O(I)$			
reaction 2	$2H_2O_2(aq) \rightarrow 2H_2O(I) + O_2(g)$			
reaction 1	$AgNO_3(aq) + HCl(aq) \rightarrow AgCl(s) + HNO_3(aq)$			

15 A student investigates the rate of reaction between magnesium and excess sulfuric acid.The volume of hydrogen given off in the reaction is measured over time.

The graph shows the results of two experiments, R and S.



Which change in conditions would cause the difference between R and S?

- **A** A catalyst is added in S.
- **B** The acid is more concentrated in R than in S.
- **C** The magnesium is less finely powdered in R than in S.
- **D** The temperature in R is lower than in S.

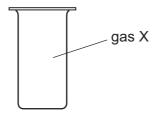
Α

16 Butane, ethanol and hydrogen are fuels.

Which substances produce **both** carbon dioxide and water when used as a fuel?

	butane	ethanol	hydrogen
Α	\checkmark	\checkmark	1
в	\checkmark	\checkmark	x
С	\checkmark	x	1
D	X	\checkmark	x

17 X is a monatomic gas.



Which statement about X is correct?

- A X burns in air.
- **B** X is coloured.
- C X is unreactive.
- **D** X will displace iodine from potassium iodide.
- **18** The equation shows the reaction between a halogen and aqueous bromide ions.

Which words correctly complete gaps 1, 2 and 3?

	1	2	3
Α	chlorine	brown	colourless
в	chlorine	colourless	brown
С	iodine	brown	colourless
D	iodine	colourless	brown

19 Carbon dioxide is an acidic oxide that reacts with aqueous calcium hydroxide.

Which type of reaction takes place?

- A decomposition
- **B** fermentation
- **C** neutralisation
- **D** oxidation
- **20** A solution contains barium ions and silver ions.

What could the anion be?

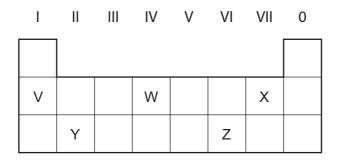
- A chloride only
- B nitrate only
- C sulfate only
- D chloride or nitrate or sulfate
- 21 A mixture containing two anions was tested and the results are shown below.

test	result
dilute nitric acid added	effervescence of a gas which turned limewater milky
dilute nitric acid added, followed by aqueous silver nitrate	yellow precipitate formed

Which anions were present?

- A carbonate and chloride
- B carbonate and iodide
- **C** sulfate and chloride
- **D** sulfate and iodide
- 22 Which is **not** a typical property of an acid?
 - A They react with alkalis producing water.
 - **B** They react with all metals producing hydrogen.
 - **C** They react with carbonates producing carbon dioxide.
 - **D** They turn litmus paper red.

23 The diagram shows a section of the Periodic Table.



Which elements will conduct electricity at room temperature?

Α	V, W and X	в	V. Y and W	С	W. X and Z	D	Y and Z
~	v, vv ana 7.		v, i ana vv	U	vv, // unu 2		i unu z

24 Water from a reservoir flows to the water works where purification processes 1 takes place followed by process 2.

What are purification processes 1 and 2?

	purification process 1	purification process 2
Α	chlorination	filtration
В	filtration	chlorination
С	fractional distillation	filtration
D	filtration	fractional distillation

25 The properties of a metal are important in deciding its use.

Which row lists a property that is **not** correct for the use given?

	use of the metal	metal property needed
Α	aluminium in aircraft wings	low density
в	aluminium in food containers	resists corrosion
С	mild steel in car bodies	high density
D	stainless steel in cutlery	does not rust

26 Brass is an alloy of copper and zinc.

Which statement is correct?

- **A** Brass can be represented by a chemical formula.
- **B** Brass is formed by a chemical reaction between copper and zinc.
- **C** The alloy will dissolve completely in dilute hydrochloric acid.
- **D** The zinc in the alloy will dissolve in dilute hydrochloric acid.
- 27 Which statement is correct for the element of proton number 19?
 - A It is a gas that dissolves in water.
 - **B** It is a hard metal that is not very reactive with water.
 - **C** It is a non-metal that burns quickly in air.
 - **D** It is a soft metal that is highly reactive with water.
- **28** Which row describes the conditions used to make steel from the iron produced by a blast furnace?

	calcium oxide (lime)	oxygen	heat						
Α	\checkmark	\checkmark	\checkmark						
в	\checkmark	\checkmark	x						
С	x	\checkmark	\checkmark						
D	X	\checkmark	X						

29 The table shows the results of adding three metals, P, Q and R, to dilute hydrochloric acid and to water.

metal	dilute hydrochloric acid	water
Р	hydrogen produced	hydrogen produced
Q	no reaction	no reaction
R	hydrogen produced	no reaction

What is the order of reactivity of the metals?

	most reactive		least reactive
Α	Р	R	Q
в	Р	Q	R
С	R	Q	Р
D	R	Р	Q

30 Which substance is a metal?

	electrical conductivity (solid)	electrical conductivity (molten)
Α	high	high
В	high	low
С	low	high
D	low	low

31 Greenhouse gases may contribute to climate change.

Two of these gases are emitted into the atmosphere as a result of processes within animals.

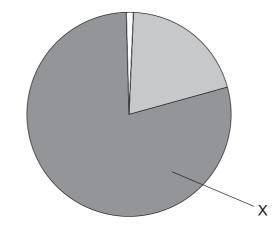
Gas1..... is produced by process3.....

Gas2..... is produced by process4......

Which words correctly complete gaps 1, 2, 3 and 4?

	1	2	3	4
Α	СО	C_2H_6	digestion	respiration
в	СО	C_2H_6	respiration	digestion
С	CO_2	CH_4	digestion	respiration
D	CO ₂	CH_4	respiration	digestion

32 The diagram shows the composition by volume of air.



What is X?

- A argon
- **B** carbon dioxide
- **C** nitrogen
- D oxygen
- **33** The table gives the composition of the atmosphere of four newly discovered planets.

planet	composition of atmosphere
W	argon, carbon dioxide and oxygen
Х	argon, nitrogen and oxygen
Y	argon, carbon dioxide and methane
Z	methane, nitrogen and oxygen

On which planets is the greenhouse effect likely to occur?

- A W only
- B W, X and Z
- **C** W and Y only
- **D** W, Y and Z

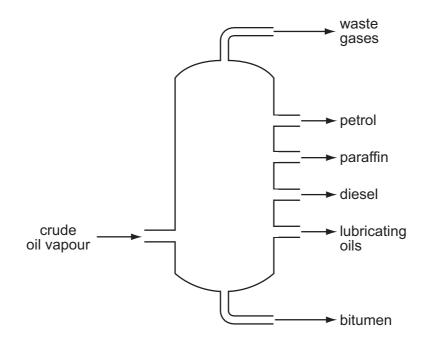
- **34** Which two substances, when reacted together, would form a salt that contains two of the essential elements provided by fertilisers?
 - A potassium hydroxide and nitric acid
 - **B** potassium hydroxide and sulfuric acid
 - C sodium hydroxide and nitric acid
 - D sodium hydroxide and sulfuric acid
- **35** Statement 1: Alloying iron with other materials to form stainless steel prevents iron from rusting by excluding oxygen.

Statement 2: Painting, oiling and electroplating are all methods of preventing iron from rusting.

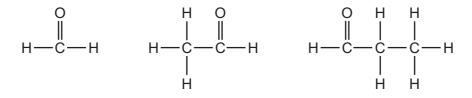
Which is correct?

- A Both statements are correct and statement 2 explains statement 1.
- **B** Both statements are correct but statement 2 does not explain statement 1.
- **C** Statement 1 is correct but statement 2 is incorrect.
- **D** Statement 2 is correct but statement 1 is incorrect.
- 36 What is the main constituent of natural gas?
 - A carbon dioxide
 - B ethane
 - C hydrogen
 - D methane
- 37 What is not essential for the formation of ethanol by fermentation?
 - A light
 - B sugar
 - C yeast
 - D water

38 Which industrial process is shown in the diagram?



- A cracking
- **B** fermentation
- **C** fractional distillation
- **D** polymerisation
- **39** The diagram shows the structures of three compounds.



Why do these three compounds belong to the same homologous series?

- A They all contain carbon, hydrogen and oxygen.
- **B** They all contain the same functional group.
- **C** They are all carbon based molecules.
- **D** They are all flammable liquids.

40 Compounds containing five carbon atoms in a molecule may have names beginning with 'pent...'.

What is the name of the compound shown?



- A pentane
- B pentanoic acid
- **C** pentanol
- D pentene

		0	4	He	Helium 2	20	Ne	Neon 10	40	Ar	Argon 18	84	Kr	Krypton 36	131	Xe	Xenon 54		Rn	Radon 86				175	Lu Lutetium 71		۲	Lawrencium 103																											
		١١٨				19	ш	Fluorine 9	35.5	CI	Chlorine 17	80	Br	Bromine 35	127	Ι	lodine 53		At	Astatine 85				173	Yb Vtterbium 70		No	Nobelium 102																											
		١٨				16	0	Oxygen 8	32	S	Sulfur 16	79	Se	Selenium 34	128	Te	Tellurium 52		Ро	Polonium 84				169	Thulium 69		Md	Mendelevium 101																											
		>				14	z	Nitrogen 7	31	٩.	Phosphorus 15	75	As	Arsenic 33	122	Sb	Antimony 51	209	Bi	Bismuth 83				167	Er Erbium 68		Еm	Fermium 100																											
		>														_																	U	Carbon 6	28	Si	Silicon 14	73	9 Ge	Germanium 32	119	Sn	50 Tin	207	Pb	Lead 82				165	Holmium 67	;		Einsteinium 99	
		≡				1	ß	Boron 5	27	٩l	Aluminium 13	70	Ga	Gallium 31	115	In	Indium 49	204	Τl	Thallium 81				162	Dysprosium 66		ç	Californium 98	bressure																										
DATA SHEET The Periodic Table of the Elements												65	Zn	Zinc 30	112	Cd	Cadmium 48	201	Hg	Mercury 80				159	Tb Terbium 65		BĶ	Berkelium 97	ature and																										
												64	Cu	Copper 29	108	Ag	Silver 47	197	Au	Gold 79				157	Gd Gadolinium 64		Cm	Curium 96	n tempera																										
DATA SHEET ic Table of th	Group											59	Ï	Nickel 28	106	Pd	Palladium 46	195	Ł	Platinum 78				152	Eu Europium 63		Am	Americium 95	m³ at roor																										
DAT/ riodic Ta						-						59	ပိ	Cobalt 27	103	Rh	Rhodium 45	192	Ir	Iridium 77				150	Samarium 62		Pu	Plutonium 94	as is 24 dı																										
The Pei			-	I	Hydrogen 1							56	Fe	lron 26	101	Ru	Ruthenium 44	190	os	Osmium 76				1	Promethium 61		Np	Neptunium 93	of any ga																										
												55	Mn	Manganese 25		ЦC	Technetium 43	186	Re	Rhenium 75				144	Neodymium 60		D	Uranium 92	one mole																										
												52	ۍ	Chromium 24	96	Mo	Molybdenum 42	184	3	Tungsten 74				141	Pr Praseodymium 59		Ра	Protactinium 91	The volume of one mole of any gas is 24 dm 3 at room temperature and pressure (r.t.p.).																										
												51	>	Vanadium 23	93	qN	Niobium 41	181	Ta	Tantalum 73				140	Cerium Cerium	232	Ч	Thorium 90	The v																										
												48	F	Titanium 22	91	Zr	Zirconium 40	178	Ħ	Hafnium 72				1		nic mass	lodi	nic) number																											
									1			45	Sc	Scandium 21	68	≻	Yttrium 39	139	La	Lanthanum 57 *	227	Ac	адпиш 89 †	series	eries	a = relative atomic mass	X = atomic symbol	b = proton (atomic) number																											
		=				6	Be	Beryllium 4	24	Mg	Magnesium 12	40	ca	Calcium 20	88	Sr	Strontium 38	137	Ba	Barium 56	226	Ra	88	*58-71 Lanthanoid series	190-103 Actinoid series	a a	×	P																											
		_				7	E	Lithium 3	23	Na	Sodium 11	39	¥	Potassium 19	85	Rb	Rubidium 37	133	Cs	Caesium 55		Ë,	Prancium 87	*58-71 L	†90-103		Key	q																											

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.