## **CAMBRIDGE INTERNATIONAL EXAMINATIONS**

**International General Certificate of Secondary Education** 

## MARK SCHEME for the May/June 2013 series

## 0620 CHEMISTRY

0620/62

Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.



	Page 2	Mark Scheme	Syllabus	Paper
		IGCSE – May/June 2013	0620	62
1	(a) pestle ar	nd / or mortar (1) filter / funnel (1)		[2]
	<b>(b) (i)</b> labe	lled arrow at liquid in mortar (1)		
	(ii) labe	lled arrow at liquid in either tube or liquid in funnel o	or any combination	n (1) [2]
	(c) (i) top l	line labelled (1)		[1]
	(ii) three	e (1)		[1]
2	(a) black (1)			[1]
	(b) (i) copp	per / Cu (1)		
	(ii) wate	er / H <sub>2</sub> O (1) <b>accept:</b> steam		[2]
	(c) boiling p	oint / freezing point (1)		
	100°C / <b>note:</b> do	0°C (1) not accept a chemical test		[2]

			IGCSE – May/June 2013	0620	62
(a)	table of results for Experiment 1				
	final and initial volumes and difference completed correctly 26.00, 0.0 and 26.0 (1)				
	to 1	deci	mal place (1) accept: volumes to 2 d.p. (e.g. 26.00	)	[2]
(b)	table of results for Experiment 2				
	final and initial volumes and difference completed correctly 19.0 and 32.0 (1) 13.0 (1) ignore: decimal places, accept: 19, 32,13, allow: ecf on final and initial volumes				
(c)	(i)		urless <b>not:</b> clear to purple / pink (1)  ept: colour change either way round		[1]
	(ii)		an acid / alkali reaction or potassium manganate cator / there is already a colour change / owtte (1)	is coloured or p	ink / acts as an [1]
(d)	(i)	expe	eriment 1 (1) allow: ecf on (a) and (b)		[1]
	(ii)	expe	eriment 1 is twice the volume of experiment 2 / exeriment 1 (1) <b>note:</b> must be a quantitative comparise table <b>allow:</b> ecf (e.g. 13 times as much as experiment 1).	on, do not allow c	
	(iii)	solu	tion B / experiment 1 more concentrated / stronger	(1) or converse	
			ole / twice (1) ore: reference to reactivity		[2]
(e)	half	value	e from table result for experiment 2 (6.5) (1) <b>allow:</b>	ecf	
	cm <sup>3</sup>	(1)			
	half	volu	me of <b>C</b> used (1)		[3]
(f)	oxidation (1) reduction (1)				
	or: electrons are lost (1) gained (1) transferred (2)  accept: oxidation numbers increase (1) decrease (1)  accept: hydrogen / H <sub>2</sub> / H lost (1) gained (1)  accept: oxygen / O <sub>2</sub> / O gained (1) lost (1)				
(g)	advantage easy to use / quick / convenient (1) ignore: large volumes				
	disa	advan	ntage not accurate / owtte (1)		[2]

Mark Scheme

Syllabus

Paper

Page 3

3

Page 4	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2013	0620	62

4 (a) colourless (1) ignore: clear, not: white [1] **(b)** white (1) precipitate (1) dissolves / clears (1) [3] (c) white precipitate (1) insoluble / does not dissolve (1) [2] (d) no change / colourless solution / no reaction (1) [1] (e) white (1) precipitate (1) [2] (g) carbon dioxide / CO<sub>2</sub>(1) [1] (h) calcium /  $Ca^{2+}$  (1) accept: any Group 2 metals carbonate /  $CO_3^{2-}$  (1) note:  $CaCO_3 = 2$ [2] 5 (a) thermometer diagrams completed correctly (3) –1 each incorrect 23, 29, 35, 41, 39, 35, 31 ignore: decimal places [3] (b) points plotted correctly (3), -1 each incorrect two intersecting straight lines (1) [4] allow: lines extending beyond intercept but must be just two lines and no curves (c)  $16 \text{ (cm}^3) \pm 0.5 \text{ (1)}$  any indication (1) [2] (d) 23(°C)(1) [1] **(e)** good insulator or reference to minimising heat losses (1) [1] (f) reaction produced heat or energy (1) accept: reaction is exothermic reaction finished / reactant(s) used up / KOH used up / neutralised (1) not: acid used up/ neutralised [2] (g) exothermic (1) [1]

Page 5	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2013	0620	62

**note:** all methods can gain the first three marks but only methods that would give usable results can gain the last three marks

known / same mass / amount of metal (1)

known / same volume / amount of acid (1)

test both **A** and **B** (1)

a method of collecting results (1)

time or run side by side (1)

comparison of results (1)

max 6 [6]