CAMBRIDGE INTERNATIONAL EXAMINATIONS

Cambridge International General Certificate of Secondary Education

MARK SCHEME for the October/November 2014 series

0620 CHEMISTRY

0620/52

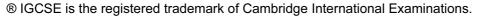
Paper 5 (Practical), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2014 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.





			Cambridge IGCSE – October/November 2014	0620	52
1	(a)	initi to 1	Table of results for Experiment 1 initial and final volumes and difference completed correctly (1) to 1 decimal place (1)		
		con	nparable to supervisors (1) ±2 cm ³		[3]
	(b)	Initi	ole of results for Experiment 2 al and final volumes completed correctly (1) I difference (1)		
			nparable to supervisors (1) ±2 cm ³		[3]
	(c)	(i)	yellow, not orange to pink / orange (1) not red		[1]
		(ii)	as an indicator / to show end point (1) ignore to see colour change		[1]
	((iii)	neutralisation (1)		[1]
	(d)	(i)	experiment 1 (1) allow: ecf from tables		[1]
		(ii)	quantitative comparison experiment 1 4X volume experiment 2/x cm³ more than (1)		[1]
	((iii)	solution B more concentrated/stronger (1) or converse explanation e.g. 4X as concentrated/less volume used (1)		[2]
	(e)	hal	f value / half value from table result for experiment 2 (1) cm ³ (1)		[2]
	(f)		vantage sy to use / quick / convenient (1)		
			advantage accurate (1)		[2]
	(g)		ne volume of each solution (1) add suitable named reactant (1) ected observation (1) comparison (1)		
			. 10 cm ³ of each acid (1) add strip of magnesium/named carbonate (ervescence (1) more rapid bubbles means stronger acid (1)	1)	[4]

Mark Scheme

Syllabus

			Cambridge IGCSE – October/November 2014	0620	52
2	(a)	(i)	purple / black / violet (1) crystals (1)		[2]
		(ii)	drops / condensation at top of tube (1) colour change to green/grey green on cooling (1)	(1)	max [2]
	(b)	(i)	green / grey (1) not white precipitate (1)		[2]
			dissolves / clears (1)		[1]
		(ii)	green / grey not white precipitate (1) insoluble (1)		[2]
	(c)		e / green (1) glowing splint (1) relights / glows brighter (1) ervescence / bubbles (1)		max [3]
	(d)	no	reaction / no precipitate / no change / colourless solution (1)		[1]
	(e)	whi	ite (1) precipitate (1)		[2]
	(f)	•	drated/water (1) w transition metal		[1]
	(g)		halide / chloride / iodide (1) sulfate (1) nsition metal / iron / chromium / catalyst (1)		[3]

Mark Scheme

Syllabus

Page 3