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**COMBINED SCIENCE**

**5129/12**

Paper 1 Multiple Choice

**May/June 2018**

**1 hour**

Additional Materials:      Multiple Choice Answer Sheet  
   Soft clean eraser  
   Soft pencil (type B or HB is recommended)



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**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

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This document consists of **16** printed pages.

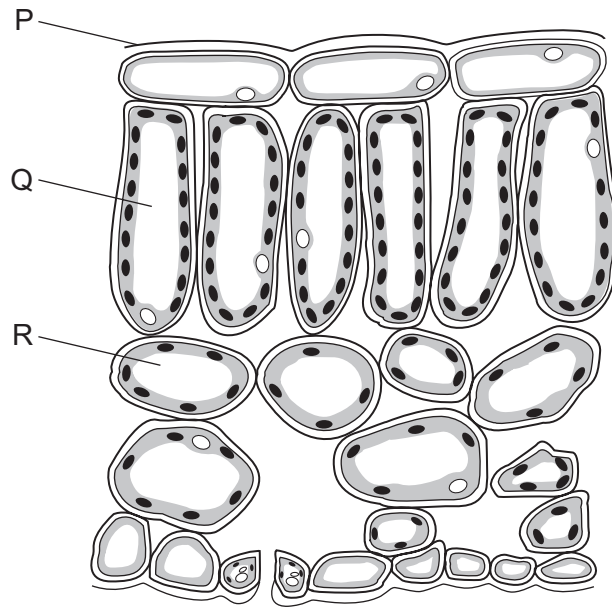
- 1 Which structures are only found in plant cells?
- A cell membrane, chloroplast and sap vacuole
  - B cell membrane, cytoplasm and nucleus
  - C cell wall, chloroplast and sap vacuole
  - D cell wall, cytoplasm and nucleus
- 2 What is the name of the process by which molecules of glucose move out of the human ileum and into the blood stream?
- A diffusion
  - B digestion
  - C peristalsis
  - D respiration
- 3 Four test-tubes contain starch solution and amylase. They are placed in water baths at different temperatures and provided with different pHs, as shown in the table.

After 30 minutes, iodine solution is added to each tube.

In which test-tube do the contents remain yellow-brown?

	temperature / °C	pH
<b>A</b>	35	2.5
<b>B</b>	35	6.9
<b>C</b>	75	2.5
<b>D</b>	75	6.9

4 The diagram shows a cross-section of a leaf.



Which row correctly identifies P, Q and R?

	P	Q	R
<b>A</b>	cuticle	palisade cell	spongy mesophyll cell
<b>B</b>	cuticle	spongy mesophyll cell	palisade cell
<b>C</b>	stomata	palisade cell	spongy mesophyll cell
<b>D</b>	stomata	spongy mesophyll cell	palisade cell

5 Which row correctly describes functions of the stomach, ileum and liver?

	stomach	ileum	liver
<b>A</b>	absorption	assimilation	ingestion
<b>B</b>	digestion	absorption	assimilation
<b>C</b>	egestion	digestion	absorption
<b>D</b>	ingestion	egestion	digestion

6 What is transpiration?

- A** absorption of water by root hairs
- B** loss of water vapour from stomata
- C** movement of water up through the xylem
- D** wilting

7 Which statement about blood capillaries is **not** correct?

- A They have thin walls.
- B They have permeable walls.
- C They have valves.
- D They supply cells with oxygen.

8 What is the function of the human alveoli?

- A breathing
- B digestion
- C gas exchange
- D respiration

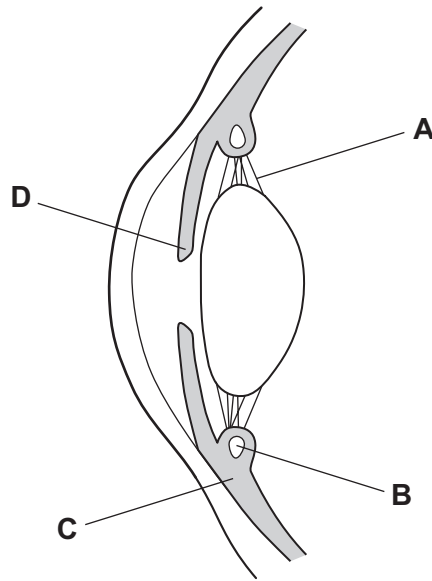
9 The body cannot store amino acids.

Which flow chart correctly shows what happens to excess amino acids in blood?

- A**    excess amino acids in the blood → broken down in kidney → urea in the urine → travel to liver → urea in the blood
- B**    excess amino acids in the blood → broken down in kidney → urea in the blood → travel to liver → urea in the urine
- C**    excess amino acids in the blood → broken down in liver → urea in the urine → travel to kidney → urea in the blood
- D**    excess amino acids in the blood → broken down in liver → urea in the blood → travel to kidney → urea in the urine

10 The diagram shows a section through part of a human eye.

Which structure contains the muscles that contract to control pupil size?



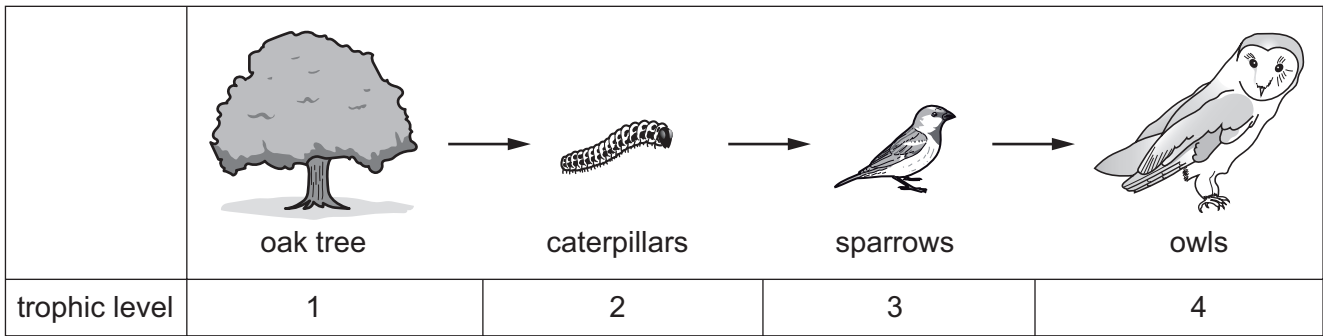
11 A student writes that taking heroin could result in four different effects.

- 1 addiction
- 2 build-up of lactic acid
- 3 stimulation
- 4 withdrawal symptoms

Which effects are correct for heroin?

- A** 1 and 2      **B** 1 and 4      **C** 2 and 3      **D** 3 and 4

12 The diagram shows a food chain.



The tree has 100 000 kJ of energy.

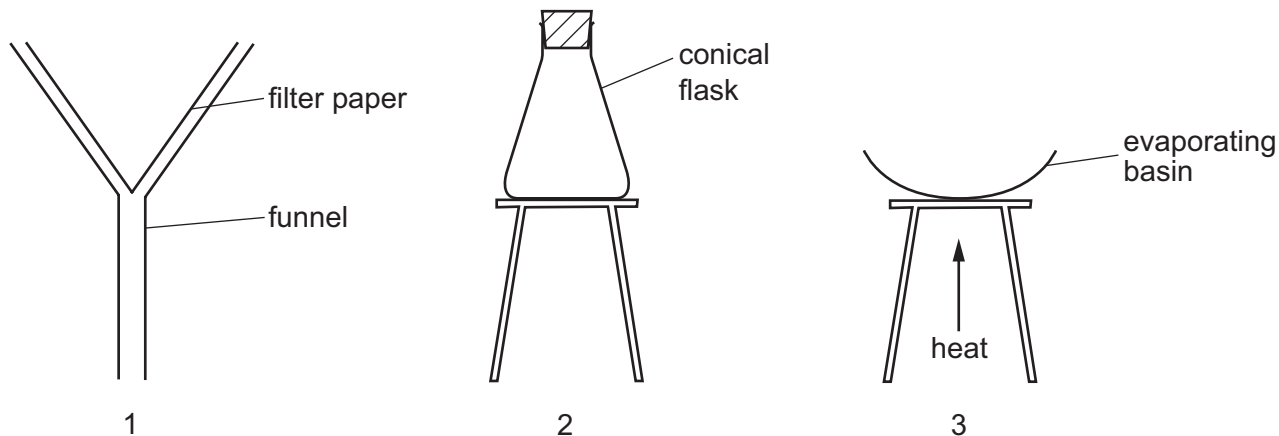
Which row indicates the likely energy transfer between each trophic level in this food chain?

	level 1 /kJ	between 1–2 /kJ	between 2–3 /kJ	between 3–4 /kJ
<b>A</b>	50	500	10 000	100 000
<b>B</b>	50	100 000	50 000	10 000
<b>C</b>	100 000	10 000	500	50
<b>D</b>	100 000	10 000	500	500

13 Which method of birth control relies on knowing the timing of ovulation?

- A** chemical
- B** mechanical
- C** natural
- D** surgical

14 The diagrams show three sets of apparatus.



Which apparatus is used to obtain separate samples of sand and salt from a mixture of sand and salt solution?

- A** 1 and 3      **B** 1 only      **C** 2 and 3      **D** 3 only

15 An atom of sodium is represented by  ${}_{11}^{23}\text{Na}$ .

What is the number of electrons in this atom?

- A** 11      **B** 12      **C** 23      **D** 34

16 Which statement about mixtures is correct?

- A** Brass is a mixture of different compounds.  
**B** Crude oil is a mixture of different elements.  
**C** Haematite is a mixture of different elements.  
**D** Wine is a mixture of different compounds.

17 Which type of bonding occurs in the compound calcium chloride?

- A** atomic  
**B** covalent  
**C** ionic  
**D** metallic

18 A compound has low electrical conductivity and high volatility.

Which type of bonding is present in the compound?

- A atomic
- B covalent
- C ionic
- D metallic

19 Aluminium oxide has the formula  $Al_2O_3$ .

What are the formulae of the aluminium ion and the oxide ion?

	aluminium ion	oxide ion
A	$Al^+$	$O^-$
B	$Al^{2+}$	$O^{2-}$
C	$Al^{2+}$	$O^{3-}$
D	$Al^{3+}$	$O^{2-}$

20 Which balanced equation for the reaction between iron and oxygen is correct?

- A  $Fe_2 + O_3 \rightarrow Fe_2O_3$
- B  $2Fe + 3O \rightarrow Fe_2O_3$
- C  $4Fe + 2O_2 \rightarrow 2Fe_2O_3$
- D  $4Fe + 3O_2 \rightarrow 2Fe_2O_3$

21 An element forms an amphoteric oxide.

Which substances will react with an amphoteric oxide?

	acids	alkalis
A	✓	✓
B	✓	x
C	x	✓
D	x	x

key

✓ = reacts

x = does not react



22 The table shows the electronic structures of five elements, V, W, X, Y and Z.

The letters are not their chemical symbols.

element	electronic structure
V	2.2
W	2.7
X	2.8.2
Y	2.8.7
Z	2.8.8

Which elements are metals in the same group?

- A** V and W      **B** V and X      **C** W and Y      **D** X, Y and Z

23 P, Q and R are three metallic elements.

The letters are not their chemical symbols.

P does not react with dilute hydrochloric acid.

Q fizzes and produces hydrogen when added to cold water.

R produces hydrogen when added to dilute hydrochloric acid but it does not react with cold water.

What are P, Q and R?

	P	Q	R
<b>A</b>	aluminium	zinc	calcium
<b>B</b>	copper	potassium	calcium
<b>C</b>	copper	sodium	zinc
<b>D</b>	zinc	potassium	iron

24 The table shows some metals and their uses.

For which metal is the correct reason given for the stated use?

	metal	use	reason
<b>A</b>	aluminium	manufacture of aeroplane wings	strength and high density
<b>B</b>	copper	electrical wiring	good conductor of heat
<b>C</b>	iron	manufacturing stainless steel	rusts
<b>D</b>	zinc	galvanising iron	zinc is more reactive than iron

25 Which substance produces hydrogen gas when it reacts with dilute hydrochloric acid?

- A magnesium
- B magnesium carbonate
- C magnesium hydroxide
- D magnesium oxide

26 Which statement about methane and petroleum is correct?

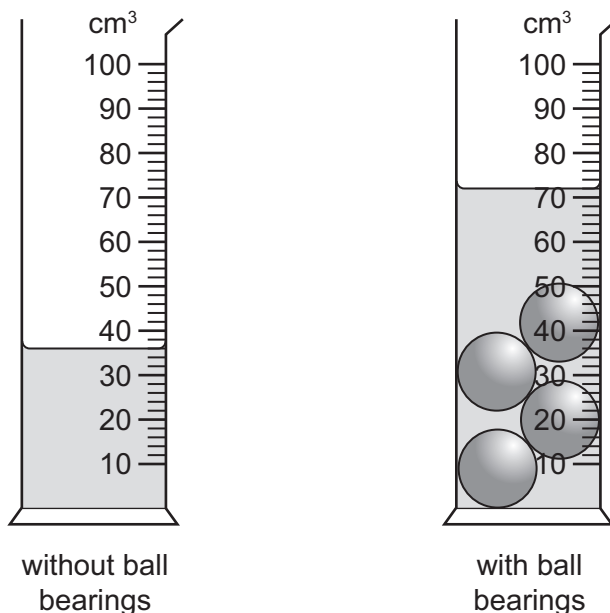
- A Exothermic reactions occur when methane and petroleum are burned.
- B Methane and petroleum are both a mixture of alkanes.
- C Methane and petroleum are both constituents of natural gas.
- D Methane has the formula  $\text{CH}_4$  and petroleum has the formula  $\text{C}_4\text{H}_{10}$ .

27 Alkenes are a series of unsaturated hydrocarbons containing a carbon to carbon double bond.

Which formula does **not** represent an alkene?

- A  $\text{C}_2\text{H}_4$
- B  $\text{C}_3\text{H}_6$
- C  $\text{C}_4\text{H}_{10}$
- D  $\text{C}_6\text{H}_{12}$

28 Four identical ball bearings are placed in a measuring cylinder containing water.



What is the volume of one ball bearing?

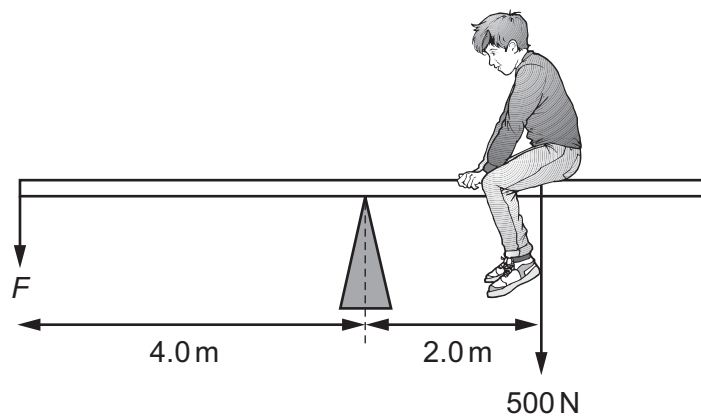
- A  $8.75 \text{ cm}^3$
- B  $9.00 \text{ cm}^3$
- C  $9.50 \text{ cm}^3$
- D  $9.75 \text{ cm}^3$

29 The gradient of the line on a graph gives the acceleration of a moving object.

What are the quantities on the horizontal and vertical axes of this graph?

	quantity on horizontal axis	quantity on vertical axis
<b>A</b>	speed	distance
<b>B</b>	speed	time
<b>C</b>	time	distance
<b>D</b>	time	speed

30 The diagram shows a boy of weight 500 N sitting on a see-saw. He sits 2.0 m from the pivot.



What is the force  $F$  needed to balance the see-saw?

- A** 250 N      **B** 750 N      **C** 1000 N      **D** 3000 N

31 Jewellery can be made from metal by first heating and then hammering the metal.

What changes during the hammering action?

- A** density  
**B** mass  
**C** shape  
**D** volume

- 32 Coal-fired power stations and hydroelectric power stations transfer a stored form of energy into other forms before producing electrical energy.

Which are the correct stored energies?

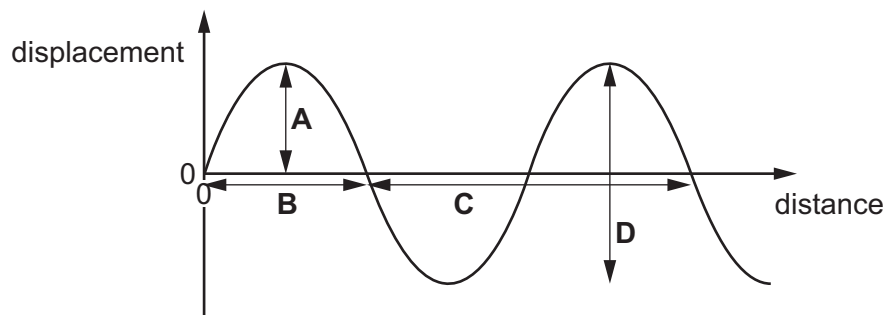
	coal-fired stored energy	hydroelectric stored energy
<b>A</b>	chemical	gravitational potential
<b>B</b>	chemical	kinetic
<b>C</b>	thermal	gravitational potential
<b>D</b>	thermal	kinetic

- 33 Which process is **not** involved when convection takes place in a gas?

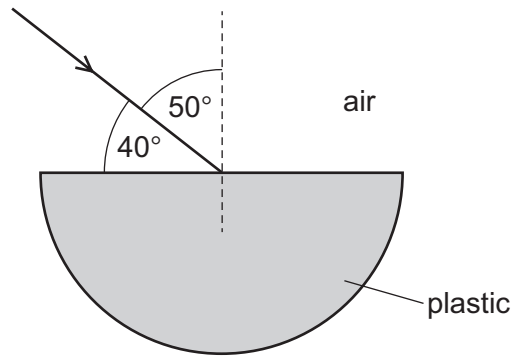
- A** The density of the gas decreases.
- B** The separation of the molecules increases.
- C** The molecules expand.
- D** The heated gas rises.

- 34 The diagram shows the displacement of the particles in a wave.

Which value is multiplied by the frequency to give the speed of the wave?



35 A ray of light is incident on the surface of a block of plastic.

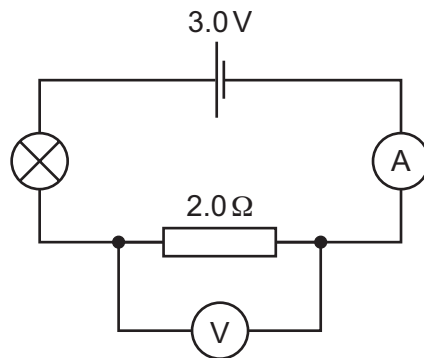


The refractive index of the plastic is 1.5.

What is the angle of refraction of the ray in the plastic?

- A**  $25^\circ$       **B**  $27^\circ$       **C**  $31^\circ$       **D**  $33^\circ$

36 The diagram shows a lamp connected in series with a resistor.

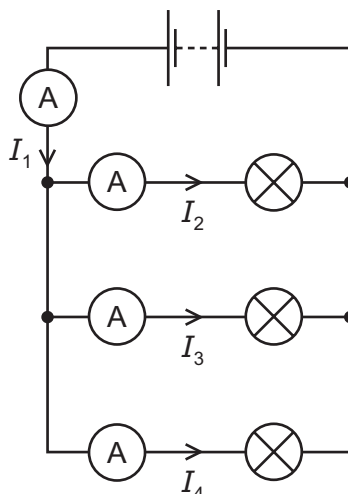


The current in the lamp is 0.5 A.

Which row gives the correct voltmeter and ammeter readings?

	voltmeter / V	ammeter / A
<b>A</b>	0.5	1.0
<b>B</b>	1.0	0.5
<b>C</b>	3.0	0.5
<b>D</b>	3.0	1.5

37 A student sets up the circuit shown.



The currents measured by the ammeters are shown.

Which equation is correct?

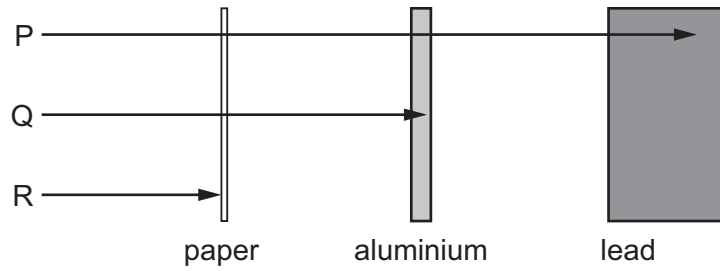
- A**  $I_1 = I_2 + I_3 + I_4$
- B**  $I_1 = I_2 = I_3 = I_4$
- C**  $I_2 + I_3 = I_4 + I_1$
- D**  $I_4 = I_3 + I_2 + I_1$
- 38 When a correctly wired plug is used to connect a working device to the mains, which statement is **not** correct?
- A** The fuse is in the live wire connection to the mains.
- B** The switch is in the live wire connection to the mains.
- C** The value of the fuse is just larger than the current.
- D** There is current in the earth wire.
- 39 The symbol for a nuclide is  ${}^{87}_{37}\text{X}$ . X represents the symbol for any chemical element.

Another nuclide has the same number of neutrons but one more proton.

What is the symbol for this other nuclide?

- A**  ${}^{87}_{36}\text{X}$       **B**  ${}^{87}_{38}\text{X}$       **C**  ${}^{88}_{37}\text{X}$       **D**  ${}^{88}_{38}\text{X}$

40 The diagram shows the penetrative powers of three types of radiation.



Which row correctly identifies each radiation?

	P	Q	R
<b>A</b>	beta	alpha	gamma
<b>B</b>	beta	gamma	alpha
<b>C</b>	gamma	alpha	beta
<b>D</b>	gamma	beta	alpha

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The Periodic Table of Elements

		Group											
I	II	III	IV	V	VI	VII	VIII						
		1 H hydrogen 1											
3 Li lithium 7	4 Be beryllium 9	<b>Key</b> atomic number atomic symbol name relative atomic mass											
11 Na sodium 23	12 Mg magnesium 24	5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20	13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84	
37 Rb rubidium 85	38 Sr strontium 88	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131	
55 Cs caesium 133	56 Ba barium 137	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —	
87 Fr francium —	88 Ra radium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	114 Fl flerovium —	116 Lv livermorium —	118 Og oganeson —	119 Uue unbinilium —	120 Uub unbinilium —	121 Uut unbinilium —	122 Uuq unbinilium —

57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).