UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS GCE Ordinary Level

MARK SCHEME for the October/November 2011 question paper for the guidance of teachers

5070 CHEMISTRY

5070/32

Paper 3 (Practical Test), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

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1 (a) Titration [12]

Accuracy 8 marks

For the two best titres give:

- 4 marks for a value within 0.2 cm³ of supervisor
- 2 marks for a value within 0.3 cm³ of supervisor
- 1 mark for a value within 0.4 cm³ of supervisor

Concordance 3 marks

Give:

- 3 marks if all the ticked values are within 0.2 cm³
- 2 marks if all the ticked values are within 0.3 cm³
- 1 mark if all the ticked values are within 0.4 cm³

Average 1 mark

Give 1 mark if the candidate calculates a correct average (error not greater than 0.05) of all his ticked values.

Assuming a 25 cm³ pipette and a titre of 20.2 cm³.

[2]

$$=\frac{25\times0.05\times2}{20.2} (1)$$

$$= 0.124(1)$$

Answers should be correct to + or -1 in the third significant figure.

(c) concentration of hydrochloric acid in scale remover

[1]

$$= 0.124 \times 10 (1)$$

$$= 1.24$$

Answer from **(b)** \times 10

(d) mass of calcium carbonate removed

[1]

$$= \frac{1.24 \times 100 \times 2}{2} (1)$$

Answer from (c) must be processed properly i.e. there must be working evident not just \times 100.

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2 R is manganese (IV) oxide, S is manganese(II) chloride, T is potassium manganate(VII)

Test			Notes
General For ppt allow so	points lid, suspension, powder		
	es f gas requires test to be at least partia ces = bubbles = gas vigorously evolve		
Solution Colourle	s ss not equivalent to clear, clear not ed	quival	ent to colourless
Solution	R		
Test 1			
	rvescence Ints a glowing splint gen	(1) (1) (1)	
Test 2			
yello	ow or brown liquid	(1)	
Test 3			
(a)	filtrate is yellow	(1)	
(b)	red-brown or brown precipitate insoluble in excess	(1) (1) (1)	
Test 4			
(a)	no reaction	(1)	
(b)	white ppt	(1)	
Test 5			
(a)	white, yellow or brown precipitate insoluble in excess colour darkens	(1) (1) (1) (1)	this mark is awarded for noting the darkening of the colour in either (a) or (b)
(b)	dark (or black) brown solid effervescence relights a glowing splint oxygen	(1) (1) (1) (1)	and dolour in citator (a) or (b)

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Test 6		
turns colourless or decolourised effervescence	(1) (1)	
Test 7		
(a) filtrate is green	(1)	
(b) filtrate turns pink, red or purple	(1)	

Conclusions

The anion in **S** is chloride or Cl⁻ (white ppt in **test 4(b)**) (1) **R** is acting as an oxidising agent (**test 2** correct or **test 3(a)** yellow or **3(b)** brown) (1) **T** is acting as an oxidising agent (decolourised or effervescence in **test 6**) (1)

Note: 25 marking points, maximum 24.